INDIAN SCHOOL MUSCAT
HALF YEARLY EXAMINATION
SEPTEMBER 2019
CLASS X

SET B

Marking Scheme – SUBJECT[CHEMISTRY][THEORY]

Q.NO.	Answers	Marks (with split up)
1.	PHY- MCQ	
2.	PHY- MCQ	
3.	PHY- MCQ	
4.	PHY- MCQ (BASED ON PRACTICAL)	
5.	a) Hydrogen gas and iron chloride is produced.	1M
6.	b) Metals are malleable and ductile.	1M
7.	b) Acidic oxides.	1M
8.	d) B and D	1M
9.	BIO- MCQ	
10.	BIO- MCQ	
11.	BIO- VSA (BASED ON PRACTICAL)	
12.	BIO- VSA (BASED ON PRACTICAL)	
13.	BIO- VSA	
14.	BIO- ASSERTION/REASONING TYPE	
15.	PHY- VSA (BASED ON PRACTICAL)	
16.	PHY- VSA	
17.	PHY- ASSERTION/REASONING TYPE	
18.	Basic , Carbon dioxide turns milky.	½ + ½ =1M
19.	Substance oxidized – CO , Substance Reduced - Fe_2O_3	½ + ½ =1M
20.	Due to evolution of CO_2 gas.	1M
21.	PHY OR	
22.	PHY	
23.	PHY OR	

24.	Cinnabar , HgS
	   i) $2HgS + 3O_2$ Heat $\rightarrow$ $2HgO + 2SO_2$
	   ii) $2HgO$ Heat $\rightarrow$ $2Hg + O_2$
	                          OR
	Diagram , Observations and conclusion
		½ + ½ = 1M
		1 + 1 = 2M
		
		3 x 1 = 3M

25.	(A) i) Endothermic ii) Exothermic (B) i) $Cu + 2AgNO_3 \rightarrow Cu(NO_3)_2 + 2Ag$ (Displacement Reaction ii) $3H_2 + N_2 \rightarrow 2NH_3$ (Combination reaction)	½ + ½ = 1M ½ x 4 = 2M

26.	(A) Alkai, Eg. NaOH or Name
	(B) Bee sting contains Methanoic acid.
	     Baking Soda being basic in nature, neutralizes acid and gives relief.
	(C) Strong – H_2SO_4 , HNO_3 (Any one from each)
	     Weak – CH_3COOH , H_2CO_3
	                          OR
		½ + ½ = 1M
		
		½ + ½ = 1M
		
		½ + ½ = 1M

		(A) Soil is acidic	½ M
		(B) i) A is $CuSO_4.5H_2O$ or (Hydrated copper suphate) B is water of crystallization. C is $CuSO_4$ or Anhydrous copper sulphate D is water ii) Any example of a hydrated salt	1/4x4= 1M
		(C) Plaster of Paris absorbs moisture and changes to gypsum, a hard mass.	½ M
			1M
27.	BIO OR		
28.	BIO		
29.	BIO		
30.	BIO		
31.	PHY OR		
32.	PHY		
33.		(A) i) By heating alone the metal oxide ii) Electrolytic Reduction/ Electrolysis of its molten ore. iii) By using suitable reducing agents like carbon (coke) or CO or by using highly reactive metals like sodium , aluminium as reducing agents.	3x1= 3M
		(B) Formation of Na_2O by transfer of electrons.	1M
		(C) One difference	1M
		OR	
		(A) Any two differences with chemical equations.	1 + 1 = 2M
		(B) Electrolytic Refining. Pure copper is used as cathode. Impure copper is used as anode. Copper sulfate as electrolyte.	½ x 4 = 2M
		(C) Correct definition with one example.	1M
34.		(A) Three balanced chemical equations.	3 x 1= 3M
		(B) Solution A has max. concentration of H+ ions. A is acidic, while B and C are Basic in nature	½ x 4 = 2M
35.	BIO OR		
36.	BIO		